NH2OH

NH_2OH

The formula arrangement suggests that OH is @ th end.

N vs O

↑ can mk 3 bonds

↖ can mk 1 bond.

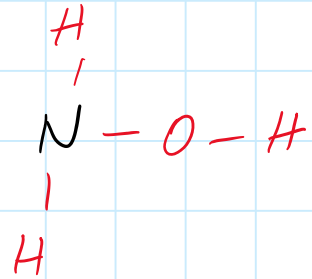
↑ goes in the middle, though the formula already seems to suggest that

$T = 5 + 2(1) + 6 + 1 = 14\bar{e}$ (no charge)

↑ N ↑ H ↑ O ↖ H

N T = 14 ē N, H, H, O, H

Follow the formula. NH_2OH

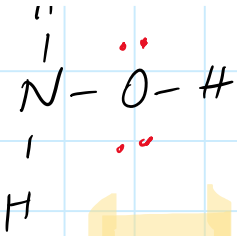


4 bonds were added ⇒ 4(2) = 8 ē used.

T = 14 − 8 = 6 ē no more atoms left

Oxygen is sort of "terminal", and need 4 ē to complete its set.

T = 6 − 4 = 2 ē

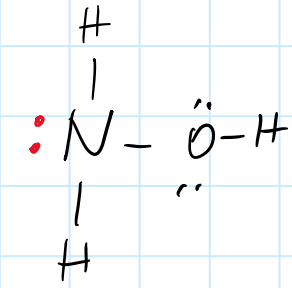


T = 6 - 4 = 2ē

This is an excellent arrangement for OH, FC = ∅

we still hv 2ē left over

N does not hv an 8-et, need 2 more ēs. ← fit together



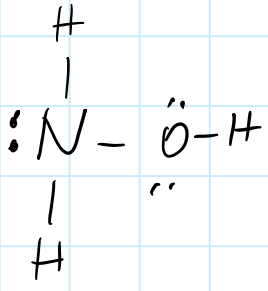
T = 2 - 2 = ∅
no more ēs left

Calculate FC-s:

FC(N) = 5 - 3 - 2 = ∅ ✓
FC(O) = 6 - 2 - 4 = ∅ ✓

Total FC-s add up to ∅, works w/ the formula being uncharged.

∴ The final structure is



you can fit the formula in other equivalent ways: