## CH3NCO

methyl isocyanate

Notice the separate pieces as suggested by the formula CH3, NCO

These should each show up in the Lewis structure.

Octets are missing for N.
when looking at which bonds to double/triple
remember that this is
methyl isocyanate
cyanide would probably be visible, and it looks
like so

Then let's try to address the bond between N and C before handling the C-O bond which has an octet, but a weak formal charge.

Is it possible there is a better arrangement that does not have so many bad Formal charges? Try something else: instead of placing electrons on Carbon first, try them on nitrogen instead:

## "the best"

b/c we removed all the negative formal charges.