

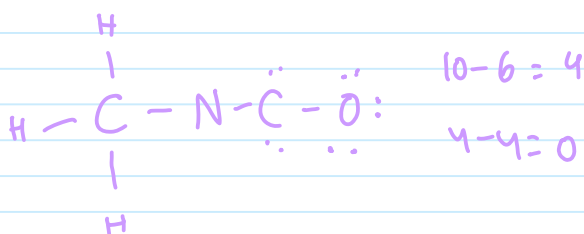
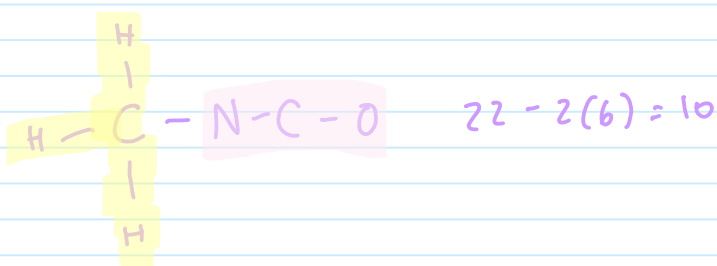
CH₃NCO

methyl isocyanate

Notice the separate pieces as suggested by the formula
CH₃, NCO

These should each show up in the Lewis structure.

$$T = 4 + 3 + 5 + 4 + 6 = 7 + 9 + 6 = 22$$



Octets are missing for N.

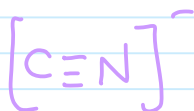
when looking at which bonds to double/triple

remember that this is

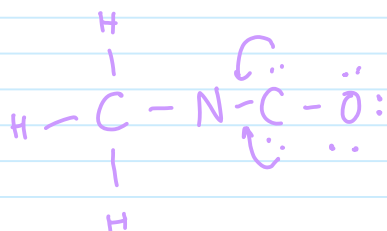
methyl isocyanate

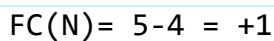
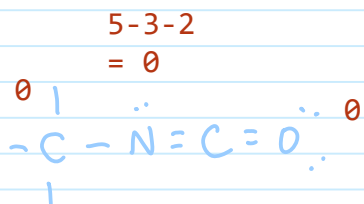
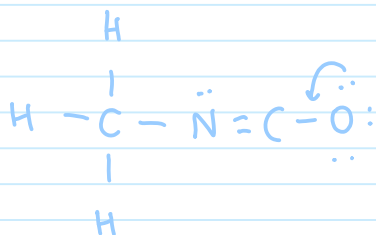
cyanide would probably be visible, and it looks like so

CN-



Then let's try to address the bond between N and C before handling the C-O bond which has an octet, but a weak formal charge.




$$\begin{array}{c} \text{H} \\ | \\ \text{H}-\text{C}-\ddot{\text{N}}-\text{C}-\ddot{\text{O}}: \\ | \quad \quad \quad \quad \quad | \\ \text{H} \quad \quad \quad \quad \quad \text{H} \end{array}$$


the best"

b/c we removed
all the negative
formal charges.

