

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
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1 1.01 1+, 1- H hydrogen	3 6.94 1+ Li lithium	4 9.01 2+ Be beryllium
11 22.99 1+ Na sodium	12 24.31 2+ Mg magnesium	
19 39.10 1+ K potassium	20 40.08 2+ Ca calcium	21 44.96 3+ Sc scandium
37 85.47 1+ Rb rubidium	38 87.62 2+ Sr strontium	39 88.91 3+ Y yttrium
55 132.91 1+ Cs cesium	56 137.33 2+ Ba barium	57 138.91 3+ La lanthanum
87 (223) 1+ 0.7 Fr francium	88 (226) 2+ 0.9 Ra radium	89 (227) 3+ 1.1 Ac actinium

Table of Common Polyatomic Ions					
acetate (ethanoate)	CH ₃ COO ⁻	chromate	CrO ₄ ²⁻	phosphate	PO ₄ ³⁻
ammonium	NH ₄ ⁺	dichromate	Cr ₂ O ₇ ²⁻	hydrogen phosphate	HPO ₄ ²⁻
benzoate	C ₆ H ₅ COO ⁻	cyanide	CN ⁻	dihydrogen phosphate	H ₂ PO ₄ ⁻
borate	BO ₃ ³⁻	hydroxide	OH ⁻	silicate	SiO ₃ ²⁻
carbide	C ₂ ²⁻	iodate	IO ₃ ⁻	sulfate	SO ₄ ²⁻
carbonate	CO ₃ ²⁻	nitrate	NO ₃ ⁻	hydrogen sulfate	HSO ₄ ⁻
hydrogen carbonate	HCO ₃ ⁻	nitrite	NO ₂ ⁻	sulfite	SO ₃ ²⁻
perchlorate	ClO ₄ ⁻	oxalate	O ₂ CCO ₂ ²⁻	hydrogen sulfite	HSO ₃ ⁻
chlorate	ClO ₃ ⁻	hydrogen oxalate	HO ₂ CCO ₂ ⁻	hydrogen sulfide	HS ⁻
chlorite	ClO ₂ ⁻	permanganate	MnO ₄ ⁻	thiocyanate	SCN ⁻
hypochlorite	OCI ⁻ or ClO ⁻	peroxide	O ₂ ²⁻	thiosulfate	S ₂ O ₃ ²⁻
		persulfide	S ₂ ²⁻		

Key		
Atomic number →	26	55.85
Electronegativity →	1.8	3+, 2+
Symbol →	Fe	
Name →	iron	

* Based on ¹²₆C
() Indicates mass of the most stable isotope

Legend for Elements			
	Metallic solids		Gases
	Non-metallic solids		Liquids

Note: The legend denotes the physical state of the elements at exactly 101.325 kPa and 298.15 K.

2 4.00 — He helium	5 10.81 2.0 B boron	6 12.01 2.6 C carbon	7 14.01 3.0 N nitrogen	8 16.00 3.4 O oxygen	9 19.00 4.0 F fluorine	10 20.18 — Ne neon
13 26.98 1.6 Al aluminium	14 28.09 1.9 Si silicon	15 30.97 2.2 P phosphorus	16 32.07 2.6 S sulfur	17 35.45 3.2 Cl chlorine	18 39.95 — Ar argon	
31 69.72 1.8 Ga gallium	32 72.64 2.0 Ge germanium	33 74.92 2.2 As arsenic	34 78.96 2.6 Se selenium	35 79.90 3.0 Br bromine	36 83.80 — Kr krypton	
49 114.82 1.8 In indium	50 118.71 2.0 Sn tin	51 121.76 2.1 Sb antimony	52 127.60 2.1 Te tellurium	53 126.90 2.7 I iodine	54 131.29 2.6 Xe xenon	
81 204.38 1.8 Tl thallium	82 207.2* 1.8 Pb lead	83 208.98 1.9 Bi bismuth	84 (209) 2.0 Po polonium	85 (210) 2.2 At astatine	86 (222) — Rn radon	

* The isotopic mix of naturally occurring lead is more variable than other elements, preventing precision to greater than tenths of a gram per mole.

lanthanide and actinide series begin

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58 140.12 1.1 Ce cerium	59 140.91 1.1 Pr praseodymium	60 144.24 1.1 Nd neodymium	61 (145) — Pm promethium	62 150.36 1.2 3+, 2+ Sm samarium	63 151.96 — 3+, 2+ Eu europium	64 157.25 1.2 3+ Gd gadolinium	65 158.93 — 3+ Tb terbium	66 162.50 1.2 3+ Dy dysprosium	67 164.93 1.2 3+ Ho holmium	68 167.26 1.2 3+ Er erbium	69 168.93 1.3 3+ Tm thulium	70 173.04 — 3+, 2+ Yb ytterbium	71 174.97 1.0 3+ Lu lutetium
90 232.04 1.3 4+ Th thorium	91 231.04 1.5 5+, 4+ Pa protactinium	92 238.03 1.7 6+, 4+ U uranium	93 (237) 1.3 5+ Np neptunium	94 (244) 1.3 4+, 6+ Pu plutonium	95 (243) — 3+, 4+ Am americium	96 (247) — 3+ Cm curium	97 (247) — 3+, 4+ Bk berkelium	98 (251) — 3+ Cf californium	99 (252) — 3+ Es einsteinium	100 (257) — 3+ Fm fermium	101 (258) — 2+, 3+ Md mendelevium	102 (259) — 2+, 3+ No nobelium	103 (262) — 3+ Lr lawrencium